

# Effect of $P_2O_5$ on Crystallization Behavior and Chemical Resistance of Dental Glasses in the $Li_2O-SiO_2-ZrO_2$ System

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**Abstract:** Commercial dental lithium disilicate based glass-ceramics containing various amounts of  $P_2O_5$  were synthesized. Regarding the crystallization behavior and physico-chemical properties of the glasses, the optimum percent of  $P_2O_5$  was determined as 8 %wt.

Crystallization behavior of the glasses was investigated by X-ray diffraction (XRD) and differential thermal analysis (DTA). The micro-hardness and chemical resistance of both glass and glass-ceramic series were also determined.

According to our results, lithium phosphate was precipitated prior to crystallization of the main phases, i.e. lithium meta silicate and lithium disilicate. This early precipitation led to evacuation of residual glass phase from lithium ions, which caused increasing the viscosity of glass and so shifting of crystallization to higher temperatures.

In addition, increasing in  $P_2O_5$  amounts and consequently increasing in  $Li_3PO_4$ , led to significant decrease in the crystallite size and aspect ratio of crystals.

Furthermore, while the chemical resistance of the glasses was decreased with  $P_2O_5$ , it was increased with  $P_2O_5$  after heat treatment process. The chemical solubility of the glass-ceramics was between 2080 and 1188  $\mu g/cm^2$ .

**Keywords:** Glass-ceramic, Lithium-disilicate,  $P_2O_5$ , Chemical solubility.

## 1. INTRODUCTION

In recent years, ceramics are used extensively in the various field of dentistry due to their biocompatibility and their similarity to tooth in points of light reflection and transmission views. Enough mechanical resistance against forces caused by chewing has been another important point for this tendency. In this regard, glass-ceramics are known as the main candidate. Today, various kinds of glass-ceramics are used in dental materials field as core, crown, opaque layer, enamels, etc.[1-6]. One of them is a commercial one named as IPSS Empress, has

been delivered to global marketing by Ivoclar Vivadent Co. This glass system is based on lithium di-silicate ( $Li_2O \cdot 2SiO_2$ ) crystalline phase with commercial name of IPS e.max [7-11].

The aim of this work was to study the effects of  $P_2O_5$  addition on the crystallization behavior, mechanical and chemical properties of a lithium di-silicate based glass, which is used by the Ivoclar Vivadent Co. leaflet as a dental core.

## 2. EXPERIMENTAL PROCEDURE

The chemical compositions of the three prepared glasses are shown in Table 1. A high

Table 1. Nominal composition of prepared samples (wt.%)

| Sample code | SiO <sub>2</sub> | Li <sub>2</sub> O | ZrO <sub>2</sub> | P <sub>2</sub> O <sub>5</sub> | Al <sub>2</sub> O <sub>3</sub> | CeO <sub>2</sub> | K <sub>2</sub> O |
|-------------|------------------|-------------------|------------------|-------------------------------|--------------------------------|------------------|------------------|
| P2          | 60               | 18                | 10               | 2                             | 3                              | 2                | 2                |
| P5          | 60               | 18                | 10               | 5                             | 3                              | 2                | 2                |
| P8          | 60               | 18                | 10               | 8                             | 3                              | 2                | 2                |

purity silica sand (>99.5%), lithium carbonate (Merck, no.1.05670), zirconium oxide (Merck, no.1.00757), aluminum oxide (Fibrona, >99%),  $P_2O_5$  (Merck, no.170540), Potassium nitrate (Merck, no.5063) and commercial cerium oxide (>99%) were used as the raw materials. After the batches were mixed completely, they were melted in alumina crucibles in an electrical furnace at 1450 °C. The heating rate was 10 °C/min, and the soaking time was 2.5 h in each experiment. The melts were then cast into steel mold, and were annealed at 500 °C for 30 minutes. After that, the specimens cooled within the kiln to room temperature, naturally.

Differential thermal analysis (DTA) was used to determine the crystallization temperature for each sample. In this way, 1.5 g of each sample with particle size between 0.3- 0.5 mm was heated up to 1000 °C with a heating rate of 10 °C/min. Heat-treatment of the glasses was performed at the determined peak crystallization temperature for one h.

The crystallinity of the samples, prior and after the heat-treatment process was identified using X-ray diffractometer (Philips PW-1800) with Cu-K $\alpha$  radiation. XRD patterns were recorded for 2 theta values from 5° to 80° with a step size of 0.1°. Before diffractometry, each sample was pulverized to reach the dimension of < 63 micro meters.

The evaluation of the glass-ceramics microstructure was done by the scanning electron

microscope (TESCAN VEGA II). In this way, the samples were mounted, polished and then chemically etched in a 2 wt. % HF solution for 25 s, then coated with a thin layer of gold.

Micro-hardness of the samples was measured using the Vickers micro-hardness tester (MXT-al) with a diamond pyramid indenter. The applied load and the loading time were 300 g and 10 s, respectively. The reported hardness number was the mean of at least ten indentations on each specimen derived by the following equation:

$$VH = \frac{1.854 F}{d^2}$$

Finally, for chemical solubility measurement of both glasses and glass-ceramics, each sample was pulverized to reach the diameter of approximately 1.895 mm. According to ISO 6872, 3g of each sample was dried at 110°C for 4 h, and then was soaked in 4vol. % acetic acid and heated at 80°C for 16 h. The lost weight of the samples was considered as a criterion for their chemical resistance[6].

### 3. RESULTS AND DISCUSSION

#### 3. 1. Differential Thermal Analysis (DTA)

Fig. 1 shows the DTA trends of the glasses. Based on the results, three exothermic peaks can be seen in each glass. According to X-ray

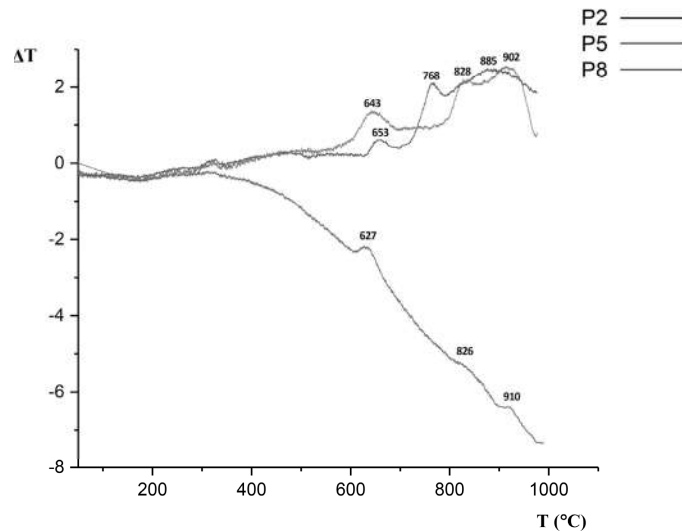


Fig. 1. DTA graph for P<sub>2</sub>, P<sub>5</sub> and P<sub>8</sub> samples with heating rate 10 °C/min

diffraction results, which are indicated in Figs. 2, 3 and 4, the first exothermic peak is attributed to crystallization of  $\text{Li}_3\text{PO}_4$ . As it can be seen, the first peak reasonably shifts to a lower temperature with increasing of  $\text{P}_2\text{O}_5$ . However, the second and the third peaks that based on the XRD results are attributed respectively to crystallization of lithium meta-silicate and lithium di-silicate, shift inversly to higher temperatures with increasing of  $\text{P}_2\text{O}_5$ . It is clear that crystallization of more lithium phosphate makes the residual glass phase more poor in

lithium ion. Therefore, it'll be reasonable that the crystallization temperature of both lithium silicate phases, i.e. meta and di-ones, shift to higher temperatures and their intensities become weaker with more using of  $\text{P}_2\text{O}_5$ . Such an effect of  $\text{P}_2\text{O}_5$  has already been reported by X. Zheng et al. [7].

### 3. 1. 1. Phase Evolution and Microstructural Survey

Based on the XRD results,  $\text{ZrO}_2$  precipitates also in the glasses with increasing of heat-

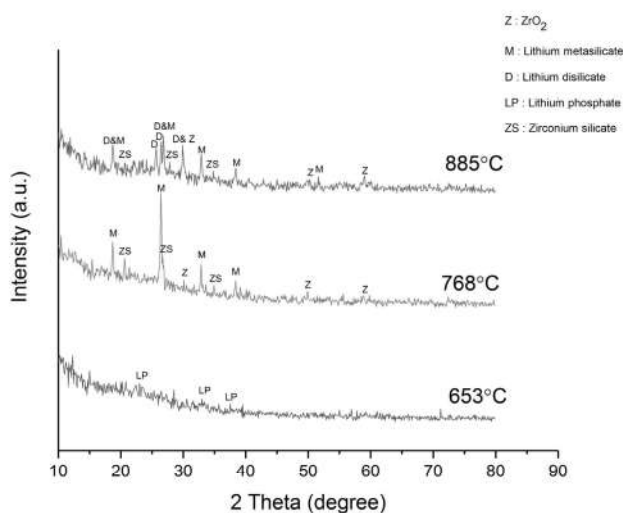


Fig. 2. XRD patterns of glass  $\text{P}_2$  after heat-treatment for 1 h at each exothermic peak temperature

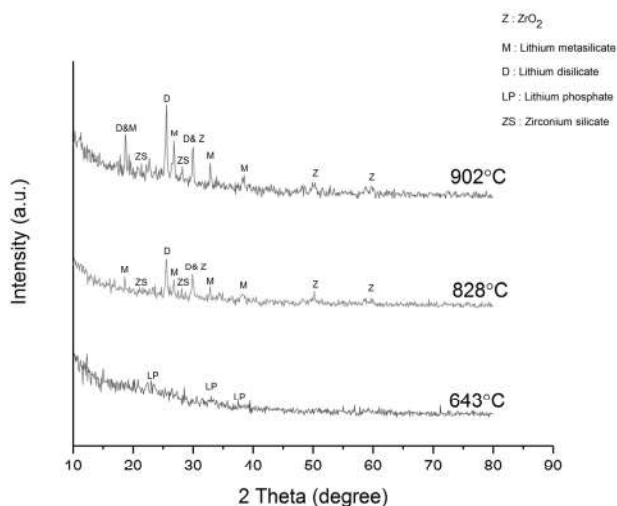


Fig. 3. XRD patterns of glass  $\text{P}_5$  after heat-treatment for 1 h at each exothermic peak temperature

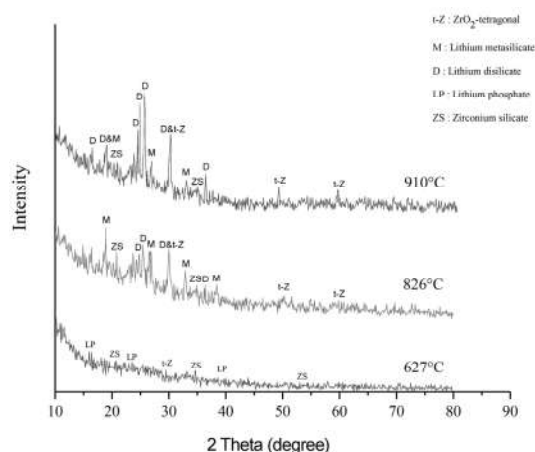


Fig. 4. XRD patterns of glass P<sub>8</sub> after heat-treatment for 1 h at each exothermic peak temperature

treatment temperature. This event is more visible for the sample P<sub>8</sub> in which zirconia shows itself as another main crystalline phase.

Fig. 5 shows the SEM micrographs of the specimens after heat-treatment at different exothermic peaks.

It is said a nucleation and growth type of liquid- liquid phase separation occurs by adding P<sub>2</sub>O<sub>5</sub>, so that amorphous phosphorous- based drops enriched in Li ions separates into a glassy matrix eventually form nano-size region [12]. During the next heat-treatment process, lithium meta-silicate precipitates which substitutes gradually with temperature by lithium di-silicate. According to our results (Figs. 5a and 5c), the morphology of the silicate phases remains unchanged during this phase alteration. It can be deduced from these results that lithium di-silicate had epitaxial growth on the lithium meta-silicate; means that lithium di-silicate has been deposited layer by layer on the meta lithium silicate substrate with the same microstructure [13-16].

By comparison of XRD patterns and SEM micrographs, it can be concluded that with increasing of P<sub>2</sub>O<sub>5</sub> a more but smaller lithium phosphate drops was formed. This will reduce the distance and/or space between the P<sub>2</sub>O<sub>5</sub> drops. This inter- drops space will then act as a spatial obstacle for growing of the main crystalline phases, i.e. lithium meta-silicate and/or lithium di-silicate, ultimately leads to finer final morphologies [13,16].

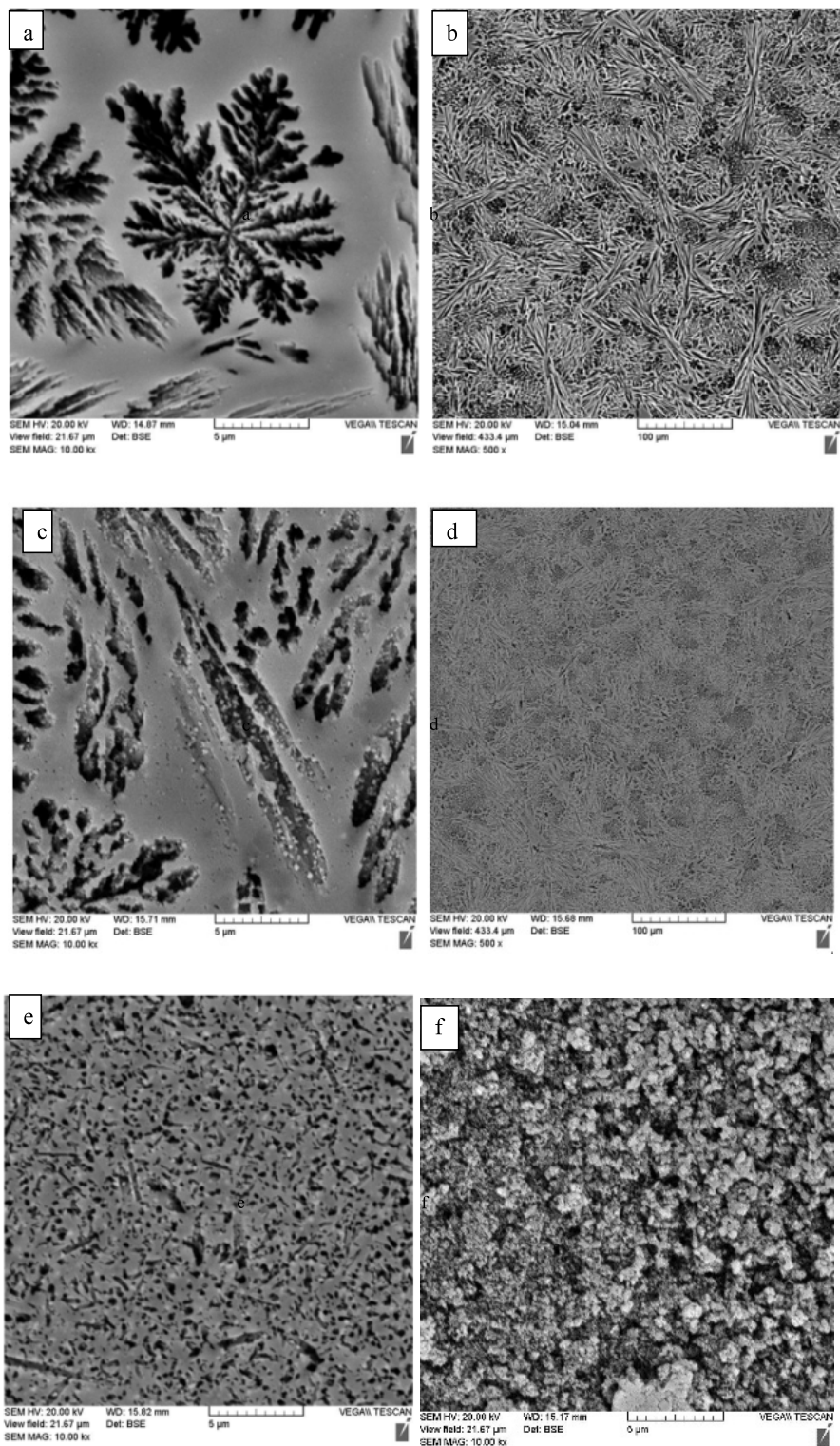
### 3. 1. 2. Vickers Micro-Hardness

Table 2 shows the average micro-hardness of at least ten indentations on each specimen.

It is known that hardness of glass-ceramics depends on the nature of both the glassy and crystalline phases, the amounts of crystalline phases, their size, their morphology, etc. In these samples, the micro-hardness of the glass-ceramics has been increased with increasing the content of P<sub>2</sub>O<sub>5</sub>. This improvement can be attributed to precipitation of much more and also finer lithium di-silicate crystals with addition of P<sub>2</sub>O<sub>5</sub>. It should also be considered that these modifications may improve other mechanical properties of the substance like strength and fracture toughness, which are essential parameters for the dental materials, act via crack deflection, crack pinning, etc. [15-18].

Table 2. Average Vickers Micro-Hardness of prepared samples

| Sample code | Vickers Micro-Hardness |
|-------------|------------------------|
| P2          | 779± 3                 |
| P5          | 995± 4                 |
| P8          | 1021± 3                |



**Fig. 5.** Microstructures of the specimens: a, b) P<sub>2</sub> after heat-treatment at its second exo-thermic peak temperature c, d) P<sub>3</sub> after heat-treatment at its third exo-thermic peak temperature e) P<sub>5</sub> after heat-treatment at its third exo-thermic peak temperature f) P<sub>8</sub> after heat-treatment at its third heat-treatment temperature

**Table 3.** Chemical solubility of the prepared samples

| Sample code | Initial weight (g) | Secondary weight after acid solubility (g) | Decreased weight (mg) | Chemical solubility ( $\mu\text{g}/\text{cm}^2$ ) |
|-------------|--------------------|--|-----------------------|---|
| GC-P2       | 3                  | 2.986                                      | 14                    | 2080.300  |
| GC-P5       | 3                  | 2.989                                      | 11                    | 1634.521  |
| GC-P8       | 3                  | 2.992                                      | 8                     | 1188.743  |
| Glass-P2    | 2.2                | 2.130                                      | 70                    | 10401.500   |
| Glass-P5    | 2.2                | 2.100                                      | 100                   | 14859.280   |
| Glass-P8    | 2.2                | 1.960                                      | 240                   | 35662.280   |

### 3. 1. 3. Chemical Solubility

If the particles were supposed to be spherical, its surface area per 3.00 g for each specimen would be approximately 672.98 mm<sup>2</sup>. Table 3. shows the results of the chemical solubility measurements of both glasses and glass-ceramics (GC).

According to standard limitations (ISIRI 12396/ISO 6872) for chemical solubility, 2000 is the maximum solubility which can be acceptable for glass-ceramics using as dental cores. Therefore, it can be concluded that except for GC-P<sub>2</sub>, the chemical solubility of the other two glass-ceramics are within the standard range. Besides, as it can be seen the chemical resistance of glass-ceramic specimens improve with increasing of P<sub>2</sub>O<sub>5</sub>. Contrary, a comparison between the three glasses shows that the least chemical resistance belongs to glass P<sub>8</sub>, whereas the highest belongs to glass P<sub>2</sub>. As it was mentioned previously, P<sub>2</sub>O<sub>5</sub> caused liquid- liquid phase separation in these glass systems[19], i.e. a phosphorous- rich and a silica- rich phases. If it happens, the P-rich region will easily dissolve in the corrosive environment leads to decrease of chemical resistance. Apparently, the solubility of Li<sub>3</sub>PO<sub>4</sub> is much less than phosphorous ions of glass; therefore, it is expected that the chemical resistance of glasses be improved by removing P<sup>+5</sup> from the glassy phase in the form of Li<sub>3</sub>PO<sub>4</sub>. Also, the volumes of residual glass phase decreases considerably with crystallization of the glass. This negligible glass phase being

surrounded by the crystalline phases which keeps away it from the access of acid solution [20-23].

## 4. CONCLUSION

The effect of P<sub>2</sub>O<sub>5</sub> on crystallization behaviors, micro-hardness and chemical resistance of three lithium silicate- based glass compositions were investigated. It was found that P<sub>2</sub>O<sub>5</sub> encouraged the crystallization of both lithium silicate phases; although it increased the crystallization temperatures of the mentioned phases. Furthermore, P<sub>2</sub>O<sub>5</sub> decreased the chemical resistance of the glasses, probably by encouraging the liquid-liquid phase separation in the glasses. On the contrary, the chemical resistance of the glass-ceramics improved by addition of phosphorous. This was attributed to precipitation of insoluble lithium phosphate phase and also protection effect of lithium silicate crystalline phases on phosphorous bearing residual glass phase.

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